30. The first three shells ($n = 1$ through 3), which can accommodate a total of $2 + 8 + 18 = 28$ electrons, are completely filled. For selenium ($Z = 34$) there are still $34 - 28 = 6$ electrons left. Two of them go to the $4s$ subshell, leaving the remaining four in the highest occupied subshell, the $4p$ subshell. Similarly, for bromine ($Z = 35$) the highest occupied subshell is also the $4p$ subshell, which contains five electrons; and for krypton ($Z = 36$) the highest occupied subshell is also the $4p$ subshell, which now accommodates six electrons.